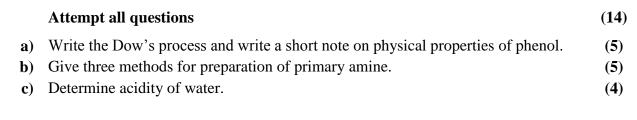
	Enrollme	ent No:		_ Exam Seat No:				
	C.U.SHAH UNIVERSITY							
	Summer Examination-2019							
	Subject Name: Chemistry-II							
	Subject Code: 4SC02CHE1		Branch: B.Sc. (All)					
	Semester	r: 2 Date:	25/04/2019	Time: 02:30 To 05:30	Marks: 70			
	(2) I (3) I	Use of Programmabl Instructions written of	on main answer land figures (if n	ny other electronic instrument book are strictly to be obeyed ecessary) at right places.	•			
Q-1	_	Attempt the follow	ving questions:			(14)		
	a) Give one example of mono and dihydric alcohol.					(1)		
	b)	-	intramolecular H-bond? Give one example.					
	c)	-						
	d)	•						
	e)	What is catalyst?				(1)		
	f)	Define: Half cell				(1)		
	g)	Give only reaction	for preparation	of Diazobenzene from anilin	e.	(1)		
	h)	Define: Hard and s	oft water			(1)		
	i)	Which species are	responsible for t	the alkalinity or basicity for v	west water?	(1)		
	j)	Define: Turbidity				(1)		
	k)	What is meant by o	=			(1)		
	1)	Which radical give	crimson red col	lor in flame test?		(1)		
	m)	What are promoter	s?			(1)		
	n)	What are amines?				(1)		
Atte	empt any f	four questions from	Q-2 to Q-8					
Q-2	2	Attempt all quest	ions			(14)		
	a)	Explain chemical p	properties of alco	ohols.		(7)		
	b)	Discuss chemical p	properties of ethe	ers.		(7)		





Q-3

Q-4	Attempt all questions		
	a)	Explain Born-Haber cycle in detail.	(7)
	b)	Write a note on non stoichiometric defects.	(7)
Q-5		Attempt all questions	
	a)	Discuss energy level diagram of H ₂ and He ₂ with bond order and its magnetic properties.	(7)
	b)	Explain the formation of σ & σ^* and π & π^* Orbitals in detail.	(7)
Q-6		Attempt all questions	
	a)	Write a note on Galvanic cell with figure.	(7)
	b)	Explain reversible and irreversible cell with proper figure.	(7)
Q-7		Attempt all questions	
	a)	Give the functions of catalyst.	(5)
	b)	Explain the working of catalyst with the adsorption theory.	(5)
	c)	Calculate ksp of Fe(OH) ₃ whose solubility 1.0×10^{-3} M.	(4)
Q-8		Attempt all questions	
	a)	Explain estimation of Ca ⁺² and Mg ⁺² with E.D.T.A. solution.	(5)
	b)	200 mL of 1.3×10^{-3} M AgNO3 is mixed with 100 ml of 4.5×10^{-5} M Na ₂ S solution will precipitations occur? ($K_{sp} = 1.6 \times 10^{-49}$)	(5)
	c)	Write a note on borax bead test.	(4)

